Species Name	Formula	$\Delta_f H°(0\ K)$	$\Delta_f H°(298.15\ K)$	Uncertainty	Units	Relative Molecular Mass	ATcT ID
Carbon monoxide	CO (g)	-113.802	-110.522	± 0.026	kJ/mol	28.01010 ± 0.00085	630-08-0*0

Contribution (%)	Reaction	Measured Quantity
63.5	C (graphite) + CO2 (g) → 2 CO (g)	$\Delta_r G°(1165\ K)$ = -33.545 ± 0.058 kJ/mol
4.3	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -393.464 ± 0.024 kJ/mol
3.6	CO2 (g) → CO (g) + O+ (g)	$\Delta_r H°(0\ K)$ = 19.0701 ± 0.0010 (×1.114) eV
3.1	CO (g) + 1/2 O2 (g) → CO2 (g)	$\Delta_r H°(303.15\ K)$ = -282.974 ± 0.116 kJ/mol
2.3	C (graphite) + CO2 (g) → 2 CO (g)	$\Delta_r G°(1236.8\ K)$ = -46.195 ± 0.3 kJ/mol
1.9	CO (g) + 1/2 O2 (g) → CO2 (g)	$\Delta_r H°(293.15\ K)$ = -283.036 ± 0.146 kJ/mol
1.8	CO (g) + 1/2 O2 (g) → CO2 (g)	$\Delta_r G°(1173\ K)$ = -180.655 ± 0.150 kJ/mol
1.7	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -393.468 ± 0.038 kJ/mol
1.7	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -393.462 ± 0.038 kJ/mol
1.2	CO2 (g) → [CO]+ (g) + O (g)	$\Delta_r H°(0\ K)$ = 19.4687 ± 0.0010 (×1.915) eV
1.1	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -94.051 ± 0.011 kcal/mol
0.8	CO (g) → C+ (g) + O (g)	$\Delta_r H°(0\ K)$ = 22.3713 ± 0.0015 eV
0.8	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -393.462 ± 0.056 kJ/mol
0.6	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(298.15\ K)$ = -393.498 ± 0.062 kJ/mol
0.6	C (graphite) + O2 (g) → CO2 (g)	$\Delta_r H°(303.15\ K)$ = -393.447 ± 0.064 kJ/mol
0.5	CO (g) + H2O (g) → CO2 (g) + H2 (g)	$\Delta_r G°(893\ K)$ = -6.369 ± 0.283 kJ/mol